Skills Worksheet

Directed Reading B

Section: Solutions of Acids and Bases STRENGTHS OF ACIDS AND BASES

concentration 1. What is the amount of acid or base dissolved in water called? **a.** concentration **b.** strength **c.** pH **d.** neutralization the acid's strength **2.** When an acid dissolves in water, which of the following is dependent on the number of molecules that break apart? **a.** the acid's concentration **b.** the acid's color **c.** the acid's durability **d.** the acid's strength a strong acid **3.** In which of the following do all the molecules of an acid break apart in water? **a.** a strong acid **b.** a strong base **c.** a weak acid **d**. a weak base a weak acic 4. In which of the following do only a few of the molecules of an acid break apart in water? **a.** a strong acid **b.** a strong base **c.** a weak acid **d.** a weak base a strong base 5. In which of the following do all the molecules of a base break apart? **a.** a strong acid **b.** a strong base **c.** a weak acid **d.** a weak base <mark>a weak base</mark> 6. In which of the following do only a few molecules of a base break apart? **a.** a strong acid **b.** a strong base **c.** a weak acid **d.** a weak base

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Name		Class	Date
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ACIDS, I	BASES, AND NEUTRALIZATION	N	
a neutralization reaction	 7. What is the reaction between a. a neutralization reaction b. an explosion c. a strength reaction d. evaporation 	n acids and bases calle	d?
water	 8. What do the H⁺ ions of an arthey react? a. oxygen b. water c. sugar d. hydrogen gas 	cid and the OH⁻ ions o	f a base form when
an indicator	 9. What can show whether a so a. an indicator b. pure water c. antacids d. salt 	olution contains an acid	d or a base?
10. The .	<mark>pH</mark> scale	e is used to express the	e acidity or basicity
(alka	linity) of a system.		
11. The ₁	pH of a solution shows the con	centration of what type	e of ion?
	H+ hydrogen ion		

Match the correct description with the correct term. Write the letter in the space provided.

7 12. pH of a neutral solution	a. greater than 7
greater than 7 13. pH of a basic solution	b. less than 7
less than 7_{14} pH of an acidic solution	c. 7
15. What are three examples of common materials v	vith a pH of less than 7?
lemon juice, acid rain, milk	

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Name	Class	Date
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16. What are three examples of commonia, bleach, sea water	ommon materials w	ith a pH of more than 7?
17. Name two types of pH indicat	cors.	
<mark>strips of pH paper and electror</mark>	lic pH meter.	
For each organism listed, write the	e preferred pH or pH ne trees	ł range.
8 - 9 19. let 20. fisl		
20. Inst 21. How does acid rain form, and Acids enter clouds from the smol this makes acid rain. Acid rain ca	what is its effect or	n nature?
and rivers causing fish to die, als		

SALTS

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22. What two substances are produced when an acid neutralizes a base?

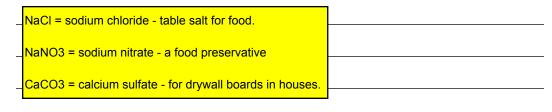
a salt and water

23. What is a salt, and how does it form?

pH of soil making plants die.

A salt is an ionic compound like NaCl,
it forms in water when an acid and
base neutralize each other.

24. Name three salts, and tell what they are used for.



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